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Sustainability of phosphorus fertilisation: sources and forms of phosphate

Sustentabilidade da fertilização fosfatada: origens e formas do fósforo

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ABSTRACT

Phosphorus (P), which originates from a non-renewable natural resource, is essential for crop productivity, food security and human health. Despite being a finite resource its use has been increasing consistently over the last few decades. The major use of the phosphate rock is the manufacture of P fertiliser, agriculture also being the main source for nonpoint source P pollution, which causes the eutrophication of water bodies. Thus, there is great need for a sustainable use of P with special attention to agricultural use, namely P fertilisation. This paper explores P sources and forms that could be phytoavailable and hence used as fertilisers. Current trends in phosphorus fertilisation are also discussed with a focus on approaches to net decrease in phosphate rock exploitation and closure of phosphorous cycle use.

Keywords: eutrophication, fertilisers, phosphate rock, soil test phosphorus, struvite

RESUMO

O fósforo (P) provém de um recurso natural não renovável e é indispensável para a produtividade agrícola, segurança alimentar e saúde humana. Apesar de ser um recurso finito o seu consumo tem vindo a crescer nas últimas décadas. O principal uso da rocha fosfatada é a produção de adubos fosfatados sendo também a atividade agrícola, a principal fonte de P de origem difusa que causa a eutrofização dos cursos de água. Estes factos enfatizam a necessidade do uso sustentável do P com especial atenção no seu uso agrícola, em particular na fertilização fosfatada. Este artigo aborda as origens e as formas de P que podem ser fitodisponíveis e, deste modo usadas como fertilizantes. São também apresentadas as atuais tendências para uma fertilização fosfatada sustentável, com referência a algumas práticas e tecnologias para alcançar não só uma real diminuição na exploração das rochas fosfatadas, com também para fechar o ciclo de utilização do fósforo.

Palavras-chave: eutrofização, fertilizantes, métodos de quantificação do fósforo, rocha fosfatada, struvite

History and evolution of phosphorus fertilisation

Phosphorus (P) is an essential nutrient for living beings making up about 0.2% of a plant's dry weight. P enters in the food chain taken up by plant roots from soil solution. P is a component of key molecules such as nucleic acids, phospholipids, and ATP. Therefore is a component of many cell constituents and plays a major role in several key processes, including photosynthesis, respiration,

and energy storage and transfer, cell division, and cell enlargement (Marschner, 1995; Schachtman *et al.*, 1998). In addition to other functions animals also require P for bones and teeth and derive their phosphorus needs from plant products and feed supplements. Thus, an adequate soil P level is essential not only for crop production but also for food security and human health.

Generally the content of the native soil available P is too low to meet crop requirements. Vance *et al.* (2003) refer that more than 40% of the world arable soils are deficient in P and needs fertilisation to increase crop productivity. Records of the use of P fertilisation in agriculture are lost in history and consist of the use of bones, "bird dung" (guano) and manures. One of the most ancient records is the use of guano in Carthage about 200 BC. The first patent of a phosphate fertiliser made by the treatment of bones with sulfuric acid was registered in 1842 by J. Lawes of Britain who subsequently modified his patent to include treatment of mineral phosphates. Liebig noted in 1857 that mineral phosphates could be solubilized with sulfuric acid as readily as bones and thus be made much more effective

as fertiliser. In only five years the production of single superphosphate (SSP) reached 200,000 tons per annum in Great Britain (Earl and Aloui, 2011).

The *Green Revolution* was a period in the mid-20th century (1930-1960) when the productivity of global agriculture increased drastically as a result of new advances and technology transfer. During this time period, new chemical fertilisers have made it possible to supply crops with extra nutrients and, therefore, increase yields. Also synthetic herbicides and pesticides were created contributing also to increase yields. Since that time the production of P mineral fertilisers uses "phosphate rock" (PR i.e. rock rich in the mineral fluorapatite $[Ca_5(PO_4)_3F]$ and phosphoric acid (H_3PO_4), which

Table 1 - Phosphate product export distribution (percent based on product tonnes - 2010)

Country	Phosphate product		
	Phosphate rock	Phosphoric acid	DAP ¹ /MAP ²
		%	
Morocco	35	48	13
Jordan	15	-	-
Tunisia	-	11	-
United States	-	8	27
China	-	-	23
Russia	-	-	16
Other countries	50	33	21

¹Di-ammonium phosphate; ²Mono-ammonium phosphate.

Source: http://www.potashcorp.com/industry_overview/2011/nutrients/25/

Worldwide uses of phosphoric acid (%)

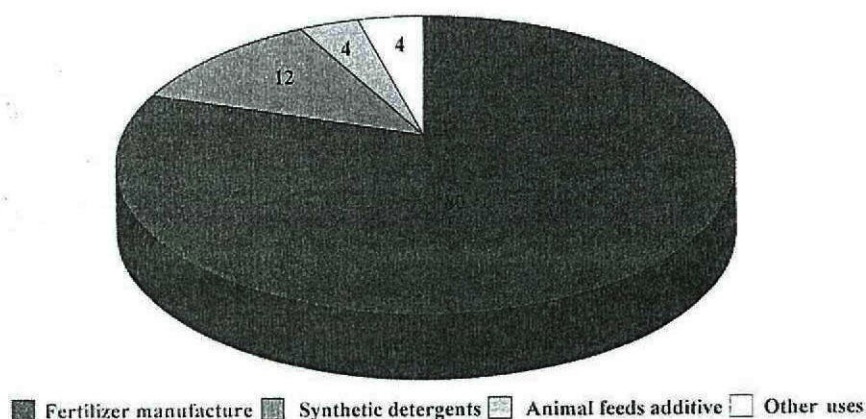


Figure 1 - Worldwide uses of phosphoric acid (%)

Source: <http://www.essentialchemicalindustry.org/chemicals/phosphoric-acid.html>

comes mainly from the treatment of PR with concentrated sulfuric acid. Single superphosphate was the first commercial P mineral fertiliser and is manufactured from the reaction between sulfuric acid and PR. Whereas triple superphosphate (TSP) fertiliser is produced by the action of concentrated phosphoric acid on ground phosphate rock. Also ammoniated phosphates (mono-ammonium phosphate-MAP and di-ammonium phosphate-DAP) are manufactured with phosphoric acid. As we can see in Table 1 Morocco is the main producer of PR and of phosphoric acid whereas US is the main producer of MAP and DAP. Data from 2013 indicates that the world annual production of phosphoric acid was about 38 million tonnes (Ciec, 2013) with 80% used for the fertiliser manufacture (Figure 1).

Nevertheless, phosphate rock is a finite natural resource and cannot be extracted forever to compensate for use losses. Most PR used in fertiliser production is sedimentary (>80%), but igneous deposits are also used and were formed mostly in shield areas and rift zones (Figure 2). Several

scarcity of PR (Runge-Metzger, 1995; Cordell *et al.*, 2009; Gilbert, 2009; Cordell *et al.*, 2011).

PR world reserves estimated by the United States Geological Survey (USGS) and by the International Fertiliser Development Center (IFDC) are quite different (Table 1). The IFDC points out that the PR world reserves are much larger than formerly estimated. Indeed some researchers have also been presenting data showing a better scenario about the concerns with PR depletion. They referred that PR reserves could meet demand into the 22nd century, and can be extended well into the 23rd century with assessed use reduction and recycling measures (Van Vuuren *et al.*, 2010; Koppelaar and Weikard, 2013). Nevertheless 95.6% of world PR reserves are located in only 5 countries having Morocco 85% of the world PR reserves (Table 2). The world distribution of PR reserves also highlights the geopolitical significance of phosphate trade. In Europe hardly no raw P is available except for very small quantities in Finland. The uncertainty about PR resource availability and consuming in the near future has led to the inclusion of PR on the EU

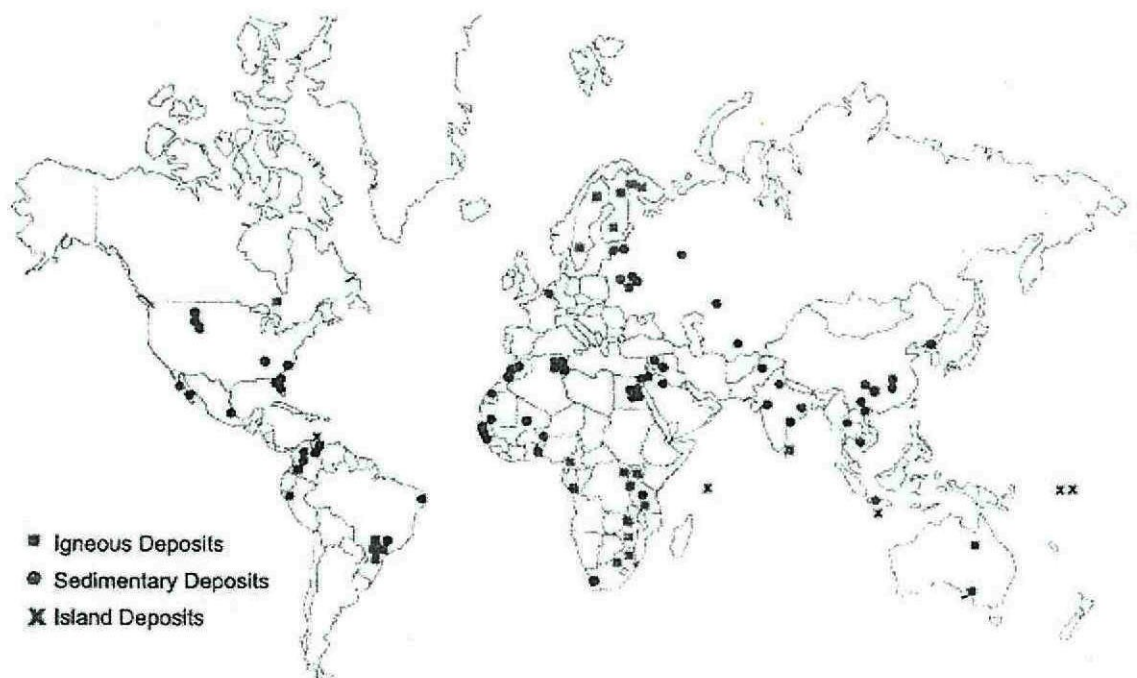


Figure 2 - Location of world phosphorus resources
Source: Zapata (2004)

researchers pointed out the concern about the fact that the current exploitation of PR is nearing the "peak production" and this shortly will drive to a

Critical Raw Materials list in 2014. Despite the risks of P resources depletion, FAO (2012) has estimated an increase of the world demand for phosphorus

Table 2 - Estimated world phosphorus rock reserves and resources (10⁹ kg)

Country	Mine production		USGS ¹ , 2010	IFDC ²	IFDC ²
	2008	2009	estimated reserves	est. reserves ³ (product)	Estimated resources ⁴
Morocco	25.0	24.0	5,700	51,000	170,000
China	50.7	55.0	3,700	3,700	16,800
United States	30.2	27.2	1,100	1,800	49,000
Jordan	6.3	6.0	1,500	900	1,800
Russia	10.4	9.0	200	500	4,300
Other countries	38.5	35.8	3,427	2,007	45,580
World total (rounded)	161	158	16,000	60,000	290,000
Estimated longevity in years based on 2008-2009 average production level			100	376	-

¹USGS – United States Geological Survey;

²IFDC – International Fertiliser Development Center;

³Reserves: materials that can be economically produced at the present time using technology;

⁴Resources: include reserves and any other materials of interest that are not reserves.

Source: International Plant Nutrition Institute [http://www.ipni.net/ipniweb/portal.nsf/0/02d5d56d777313b2062577ce0069a3a8/\\$file/p%20fert%20tech%2011%2010%202010.pdf](http://www.ipni.net/ipniweb/portal.nsf/0/02d5d56d777313b2062577ce0069a3a8/$file/p%20fert%20tech%2011%2010%202010.pdf)

fertiliser in the period 2012-2016 (from 41,525 thousand tonnes in 2012 to 45,013 in 2016), with an annual growth rate of 2% but when considering only Europe this rate increases to 2.5%. The annual growth rate for the same period in Europe for N and K was 1.3% and 2% respectively.

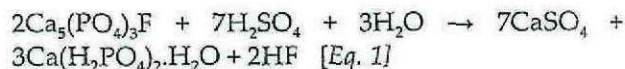
Agricultural P uses above plant needs, mainly in developed countries, has resulted in the accumulation of P in soils with resulting P losses to drainage and runoff waters and sediments, leading to the eutrophication of water bodies (Sharpley and Tunney, 2000; Carpenter, 2005; Kleinman *et al.*, 2011). Therefore sustainability of P fertilisation needs the implementation of innovative technological approaches on P and crop productivity management.

Sources of phosphorus

Mineral fertilisers from phosphate rocks

Since the middle of 19th century the P fertilisers commonly used have been the superphosphates and according to their P level they are called single (SSP) or triple (TSP).

SSP contains 7 to 9% of P (16 to 20% P₂O₅), 18 to 21% of Ca and 11 to 12% of S, and results from the reaction of PR with sulfuric acid:



TSP results from the reaction of PR with phosphoric acid, and contains ≈19% of P (45% P₂O₅) with ≤ 3% of S and 15% of Ca. The TSP chemical formula is Ca(H₂PO₄)₂·H₂O.

The production of superphosphates has declined since the 1980s in favour of ammoniated phosphates as costs of producing TSP can be higher than ammoniated phosphates, making the economics for TSP less favourable in some situations. These fertilisers could be di-ammonium (DAP: 18-46-0) or mono-ammonium phosphates (MAP: 11-52-0). Ammoniated phosphates are formed when a solution of phosphoric acid is added to ammonia until the solution is distinctly acidic. MAP has the chemical formula of NH₄H₂PO₄ and DAP is (NH₄)₂HPO₄. DAP is currently the major P dry fertiliser used. There are also in the market NPK fertilisers obtained by adding K salts to DAP, MAP, SSP or STP. NPK fertilisers can also be obtained from nitrophosphates which is originated after treatment of PR with HNO₃. There are also in the market P fertilisers with soluble P forms and liquid P solutions.

Organic wastes as P sources

Animal manures and other agro-wastes as well as treated wastewaters, sewage sludge, anaerobic digestates from biogas production plants are

Table 3 - Phosphorus and nitrogen contents (%) and N/P ratio of different organic wastes used as fertilisers in the Portuguese region of "Beira Interior Centro"

	Sewage sludge (n = 42)			Cattle slurry (n = 7)		
	P ₂ O ₅	N	N/P	P ₂ O ₅	N	N/P
Mean	1.67	3.79	7	0.034	2.85	2471
Median	1.64	3.52	5	0.010	3.24	666
Maximum	5.14	5.07	42	0.160	6.22	7133
Minimum	0.17	1.47	2	0.002	0.09	6
SD ¹	0.90	1.41	7	0.062	2.59	3022
CV ² (%)	54	37	100	102	91	122

¹Standard deviation;

²Coefficient of variation = [(SD/mean) × 100]

Source: Horta and Torrent (2010)

common organic wastes used in agricultural soil with fertiliser value. These organic wastes contain different levels and forms of phosphorus which lead to also different agronomic effects. Tables 3 and 4 shows P and N content of some organic wastes and composts obtained in Portuguese region "Beira Interior Centro" and Table 5 shows the P and N content indicated in the Portuguese Book of Best Agricultural Practices (BAP) for a rationale nitrogen planning fertilisation.

P content and N/P ratio in organic wastes and composts differs greatly not only between them but particularly in relation to the same organic waste or compost (Tables 3 and 4). These feature leads to some constraints in fertilisation strategy. If the amount of the organic waste application is calculated to meet the existing BAP to prevent

Table 4 - Phosphorus and nitrogen contents (%) and N/P ratio of different composts used as fertilisers in the Portuguese region of "Beira Interior Centro"

	Compost SS ³ (n = 5)			Cattle manure compost (n = 4)			Compost MSW ⁴ (n = 5)		
	P ₂ O ₅	N	N/P	P ₂ O ₅	N	N/P	P ₂ O ₅	N	N/P
Mean	0.66	1.61	8	1.39	2.76	12	0.28	1.77	18
Median	0.57	1.76	5	-	-	9	0.29	1.71	14
Maximum	1.38	2.08	14	2.45	3.99	28	0.36	2.1	37
Minimum	0.30	0.94	3	0.33	1.53	1	0.10	1.6	12
SD ¹	0.38	0.38	5	1.49	1.74	10	0.10	0.20	10
CV ² (%)	58	24	60	108	63	83	37	16	54

¹Standard deviation;

²Coefficient of variation = [(SD/mean) × 100];

³SS - Compost from sewage sludge and sawdust;

⁴MSW - Compost from municipal solid wastes.

Source: Horta and Torrent (2010)

Table 5 - Phosphorus and nitrogen levels of different organic wastes and composts used as fertilisers referred on Portuguese Book of Best Agricultural Practices

Organic residue or Compost	P ₂ O ₅ total	N total
Sewage sludge (%)	4.3	3.3
Cattle slurry kg m ⁻³	1.2 - 2.1	5.0 - 3.7
Swine slurry kg m ⁻³	3.5	6.0
Poultry manure kg/ton	18 - 25	40 - 20
Horse manure kg/ton	2.5	5.0
Manure of sheep and goats kg/ton	2.5	5.5
Compost MSW (%)	1.4	1.8

¹Source: Ministério da Agricultura, do Desenvolvimento Rural e das Pescas (1997)

nitrogen losses, in some cases there are a marked unbalance in N/P ratio e.g. N/P = 2 which is quite different from the N/P ratio to meet crop needs (N/P = 7-11). This generally accounts for an accumulation of P in soil. On the other hand if the amount of residue is estimated to meet the P needs of crops it is possible in some cases to apply N in excess e.g. N/P = 42. Therefore, when using organic wastes or composts as fertilisers is advisable to know its own chemical composition to do an adequate fertilisation planning instead of using published data.

Fertilisers manufactured from organic wastes

The inclusion of P in the EU Critical Raw Materials list in 2014 pointed to the need of recover and reuse P in wastes namely in food wastes, in wastewater and sludge and in manures or livestock effluents. Regarding to mineral phosphorus requirements Europe has a high import dependency above 90% of its needs. In Switzerland for example, a quantity equalling of 15% of Europe's mineral P demand has being wasted as disposed sewage sludge and its ashes (P-REX policy brief in Scope newsletter, March 2015). There are currently technological processes to remove nitrogen and phosphorus from wastewater by precipitating them as the mineral struvite (magnesium ammonium phosphate - NH₄MgPO₄·6H₂O) which contains 5% of N, 28% of P₂O₅ and 10% of Mg. Struvite was firstly known to be a "problem" material by wastewater operators. Frequently it is formed in sewage sludge digesters and liquid manure handling systems, where it causes blockages (Doyle and Parson, 2002; Mohan *et al.*, 2011), but nowadays it becomes a product that could be of economic interest. The good agronomic properties of struvite are known since the middle of 20th century (Bridger *et al.*, 1961). The use of technology to separate nutrients from wastewater, sewage sludge and its ashes and after

that use these mineral nutrients to make fertilisers will have the benefit to overcome limitations with diffuse contaminants, such as pharmaceuticals and organic chemicals arising from the direct application of sewage sludge and ashes on arable land. There are approximately 10 companies/water utilities (legal entities) producing recovered struvite from wastewater today in Europe (scope newsletter, March 2015). In addition, a proposal for "EU fertiliser criteria for recovered struvite" to include in the EU Fertiliser Regulation revision and/or struvite End-of-Waste criteria is now in discussion and validation (Proposed EU fertiliser regulation criteria for recovered struvite, 2015).

A further way of reusing nutrients has been the production of organic fertilisers after treatment of manures. The *Flemish Coordination Centre for Manure Processing* is an example of P recycled from manures. In this case the manure of pig, poultry and cattle manure or digestates undergoes liquid/solid separation, drying, composting, nitrogen removal or ammonia recovery and post treatment (e.g. addition of other nutrients, granulation) to produce an organic fertiliser, adapted for use on arable crops, so recycling both phosphorus and organic carbon. Flanders exports manures to North France and also exports granulated poultry manure to Asia and the Middle East. In total some 24,000 tonnes of phosphorus (P) are exported and recycled annually (Scope newsletter, March 2015).

These are some innovative strategies to improve the sustainability of P fertilisation contributing to close the cycle of P use.

Phosphorus forms and P phytoavailability in fertilisers

Regardless of the P sources, fertilisers should have P in available forms for plant uptake. Availability depends not only of P soil dynamics and of crop uptake characteristics but firstly it depends on the dissolution kinetics of the chemical forms of P in the fertilisers. P forms in the fertilisers that are soluble in water are in principle high available to plants. Nevertheless, the soluble P in soil solution could undergo rapid immobilisation by sorption onto soil mineral surfaces (oxi-hydroxides of Fe and Al, clay minerals and calcite), or by incorporation into soil metal-organic matter complexes (Gerke and Hermann, 1992; Matar *et al.*, 1992; Frossard *et al.*, 1995; Torrent, 1997; Borggaard *et al.*, 2004). So, the concentration of orthophosphate-P in the soil

solution is basically controlled by adsorption/desorption and precipitation/dissolution processes and also by the immobilisation and mineralisation of organic P forms. Labile P forms are the amount of P in the soil solid phase that will readily resupply P in soil solution. Single superphosphate [$\text{Ca}(\text{H}_2\text{PO}_4)_2 \cdot \text{H}_2\text{O}$, $\text{CaSO}_4 \cdot n\text{H}_2\text{O}$ ($n=0; 0.5$ or 2.0)] and triple superphosphate [$\text{Ca}(\text{H}_2\text{PO}_4)_2 \cdot \text{H}_2\text{O}$] have 85 and 87% of P forms water soluble respectively. Mono-ammonium phosphate (MAP - $\text{NH}_4\text{H}_2\text{PO}_4$) have a lower P water solubility of 50% as it can content re-precipitated apatite [$(\text{NH}_4\text{H}_2\text{PO}_4, \text{CaHPO}_4, (\text{NH}_4)_2\text{HPO}_4$)] but in di-ammonium phosphate [DAP - $(\text{NH}_4)_2\text{HPO}_4$] P forms are 100% water soluble (Engelstad and Terman, 1980). In P mineral fertilisers the standard method to evaluate P availability is by the solubility of P in ammonium citrate solution (AOAC, 1975). The solubility of P in 2% citric acid is also used to evaluate fertiliser value of basic slag a by-product of steelmaking (<2% of P solubility in water, Ca-silico-carnotite) or PR (<1% of P solubility in water, carbonate apatite - [$\text{Ca}_5(\text{PO}_4)_3(\text{OH}, \text{F}, \text{Cl})$] (Engelstad and Terman, 1980). P forms that are soluble in citric acid accounts for slow release of P from the fertilisers.

In composts and livestock manures P forms differs widely not only between different organic products but also between the same types of product. Nevertheless in these organic products P is mainly in inorganic form (Pi) and accounts approximately for 90% of total P. Frossard *et al.* (2002) evaluated Pi forms by a sequential extraction (Traoré *et al.*, 1999) and by isotopic exchange experiments in 16 composts obtained in Switzerland mainly from solid urban organic wastes (SUOW), woody wastes (WW) and farmyard manure plus SUOW or farmyard manure plus WW. The results showed that the composts contain relative large concentrations of rapidly available Pi and that a fraction of the slowly or non-exchangeable Pi is bound to Ca in the form of apatite or octacalcium phosphates. These authors concluded that from an agronomic point of view the fraction bound to Ca of compost-P will probably have a low availability to plants in neutral and alkaline soils.

Also, in six organic residues produced in Portugal and evaluated by the same P sequential extraction Pi ranges from 80 to 99% of total P (Horta, 2015). The organic residues under study were four composts obtained from the mixture of sewage sludge plus saw dust and two dry fractions from poultry or

pig slurries. In these composts the majority of Pi is in highly available forms whilst in the livestock slurries the main forms are Pi-Ca species which could have lower availability. The partition of Pi forms in struvite and TSP by the same sequential procedure was also done. The results show that 32% and 86% of readily available forms of Pi are present in struvite and TSP respectively and that the slowly available Pi forms accounts for 26% in struvite and only 13% in TSP. These results suggest that the partition of Pi forms in composts, animal manures and mineral fertilisers will influence the agronomic efficiency of P. The phytoavailability of these different P forms will depend also on soil properties (e.g. pH) and crop uptake characteristics.

Nevertheless, fertilisation of ryegrass (*Lolium perenne* L.) with animal manures (cattle manure compost, slurry dry fractions one of poultry and another from pig) or P mineral fertiliser (TSP) using the same amount of P in fertilisation (50 kg P ha⁻¹) showed that animal manures were a source of phytoavailable P (Horta *et al.*, 2014a). In addition higher yields were obtained when using animal manures as a P source compared with equivalent amounts of mineral fertiliser (Reddy *et al.* 1999, 2000; Whalen *et al.*, 2000; Horta *et al.*, 2014a). Experiments with struvite show that it's a source of P to crops with a P slow release effect which is a benefit for pastures (Thompson *et al.*, 2012; Katanda *et al.*, 2013; Cerrillo *et al.*, 2015). Nevertheless to use struvite as a fertiliser is still needed more research to develop efficiency and scale of recovered struvite from sewage and to formulate appropriate fertilisers (P-REX policy brief, March 2015).

Current phosphorus fertilisation strategy

Until now P fertilisation has been done considering the information provided by soil P tests with agronomic meaning and thus based on fertility classes or indexes associated to the ranges in the amount of soil available P (e.g. Bray, 1945; Olsen *et al.*, 1954; Egnér *et al.*, 1960; Mehlich, 1984). The amount of P extracted for each of these tests is correlated with P uptake by crops so it means in principle that the amount of P quantified by the test is phytoavailable. The currently overfertilisation observed in the majority of arable land in developed countries should have multiple sources.

First the suitability of P tests to soil properties and soil P dynamics. The amount of P in solution is the result of equilibrium between the activity

of soil solution P (intensity factor) and P sorbed onto the soil solid phase (capacity factor). Some tests are closer to quantify the capacity factor and others the intensity factor. Therefore a chemical extraction solution has not the ability of put together these dynamic relationships resulting information that should be regarded together with some other soil properties and even with crop uptake characteristics (Horta and Torrent, 2010).

A second feature is the suitability of the referred P tests concerning the soils where fertilisation was done with manures and others organic wastes. After the addition of these fertilisers the partition of P by different soil P forms reflects the composition of the organic wastes and the soil properties (Sharpley *et al.*, 2004). The chemical solutions used by soil P tests could not have, in these soils, a good correlation with crop P uptake (Horta *et al.*, 2014b) and therefore they have limitations for P fertiliser planning. As said previously in many countries in Europe the restriction in applying manures to the soil is based on its N content which have in some cases N/P ratios lower than the crop needs leading to P soil accumulation.

In Portugal as well as in other countries soil P status is characterised by a high level in intensive agricultural systems and a P deficit in extensive agricultural systems (Torrent *et al.*, 2007; Horta and Torrent, 2010). The higher levels of soil P did not result in an increase in P agronomic efficiency, FAO (2006) refer that only 15% to 30% of P in fertiliser is taken up by crops, often because fertilisers are applied far in excess of plant demand. Childers *et al.* (2011) refer that only about 20% of all mined P is actually consumed in food, the remaining 80% is lost from the human P cycle. The overfertilised soils have exceeded its P sorption capacity and behave as "hotspots" contributing to nonpoint source pollution leading to long-term eutrophication of water bodies (Heckcrath *et al.*, 1995; McDowell *et al.*, 2015). However even the extensive agricultural systems (e.g. pastures, outdoor livestock production) characterised by low inputs of production factors such as mineral fertilisers could have also "hotspots" with high soil P levels due to continuous input of feed and animal excretions (Salomon *et al.*, 2007; Horta *et al.*, 2012). These areas are also net contributors to diffuse P emissions at catchment scale. There are also in some countries e.g. Portugal, fertilisation recommendations target to the different crops (e.g. cereals, horticultural

crops, pastures, ornamentals) based on field experiments. Even these recommendations have sometimes an overestimation of P fertilisation. In addition to soil P status and crop needs it should also be considered the edaphic conditions as soil properties, topographic location and climate for an adequate fertiliser planning.

Trends in phosphorus fertilisation

Sustainability of P fertilisation will require a paradigm shift in the current P and crop productivity management. Therefore to achieve P sustainability not only a net decrease in PR exploitation is needed but it is also crucial to close the cycle of phosphorus use. In order to meet this objective Withers *et al.* (2015) have articulated five key R strategies (5R) to improve P management: Realign P inputs more precisely to maximize efficiency, Reduce P losses to the oceans, Recycle more P in bioresources, Recover and reuse P from wastes and Redefine P requirements in the food chain.

In order to Realign P inputs it has been paid some attention to the role and significance of soil organic matter and microorganisms in promoting better soil conditions for healthy plants and better nutrient uptake. So, nowadays the addition of humic substances, seaweed and beneficial microorganisms to some compound fertilisers are also done. These fertilisers provide an increase in crop uptake of nutrients by stimulating the crop metabolism and also the microbial activity on the mineralisation of soil organic-P compounds. Also products with only beneficial microorganisms such as *Azotobacter*, *Azospirillum*, *Bacillus*, *Pseudomonas*, *Rhizobium* and some genus of *Cyanobacteria* are also in the market. Mycorrhizae a mutualistic relationship between plant roots and fungus of the genus *Ascomycetes*, *Basidiomycetes* and *Glomeromycota*, improve nutrient crop uptake, mainly soil P. The increase in soil P uptake occurs due to phosphatase excretion by the fungus promoting organic P mineralisation and also by the increase in the volume of scavenged by the fungus bypassing Pi depletion in rhizosphere (Smith and Read, 1997; Smith *et al.*, 2011). Thus, some research trends are about soil inoculation with hyphae of mycorrhizal fungus or the use of seeds with the addition of phosphatase enzymes. Nevertheless the increase in better biotic conditions in soils improving agricultural practices to increase the organic matter in soil and promote the development

of soil microorganisms is a cycle that will enhance soil fertility and P uptake saving P fertilisers.

Recover and reuse P from wastes namely to close the cycle of anthropogenic P use could be achieved either by a "traditional approach" and in this case the objective is to reuse the P added to organic wastes (e.g. sewage sludge, livestock slurries, digestates, ashes) by the use of these products as fertilisers or by an innovative approach using the organic residues to obtain by technological processes some nutrients such as N, P, K and Ca. In this case struvite is a good example because there are now some enterprises in Europe producing struvite from sewage sludge and other wastes. This product could be considered both as a fertiliser (5% N, 28% P₂O₅ and 10% Mg) and either as a raw material (ingredient) for further fertiliser production. This innovative perspective changes the approach to P sustainability as it creates a new market aiming to produce safe fertilisers from organic wastes that are produced locally but can be used globally. Other R strategies concerns about exploitation of legacy P. Legacy P is the P that has accumulated in soils from past inputs of fertilisers and manures. These P forms in arable soils are a secondary source of P for crop uptake. So, legacy P could substitute manufactured fertilisers help to preserve critical reserves of finite phosphate rock and its uptake will gradually improve water quality (Rowe *et al.*, 2015). Also, plant breeding to achieve a better use of soil P and improvement of technologies as precision farming are now some trends to get sustainability of P fertilisation.

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